

KEY

Naming Ionic Compounds Worksheet One

Give the name of the following ionic compounds:

- 1) Na_2CO_3 Sodium Carbonate
- 2) NaOH Sodium hydroxide
- 3) MgBr_2 Magnesium Bromide
- 4) KCl Potassium chloride
- 5) FeCl_2 Iron (II) chloride
- 6) FeCl_3 Iron (III) chloride
- 7) Zn(OH)_2 Zinc hydroxide
- 8) BeSO_4 Beryllium Sulfate
- 9) CrF_2 Chromium (II) Fluoride
- 10) Al_2S_3 Aluminum sulfide
- 11) PbO Lead (II) oxide
- 12) Li_3PO_4 Lithium Phosphate
- 13) TiI_4 Titanium (IV) Iodide
- 14) Co_3N_2 Cobalt (II) Nitride
- 15) Mg_3P_2 Magnesium Phosphide
- 16) $\text{Ga(NO}_2)_3$ Gallium (III) Nitrite
- 17) Ag_2SO_3 Silver sulfite
- 18) NH_4OH Ammonium hydroxide
- 19) Al(CN)_3 Aluminum cyanide
- 20) $\text{Be(CH}_3\text{COO)}_2$ Beryllium acetate

For the following compounds, give the formulas

- 22) sodium phosphide $\overset{+1}{\text{Na}}_3 \overset{-3}{\text{P}}$
- 23) magnesium nitrate $\overset{+2}{\text{Mg}} (\overset{-1}{\text{NO}_3})_2$
- 24) lead (II) sulfite $\overset{+2}{\text{Pb}} \overset{-2}{\text{SO}_3}$
- 25) calcium phosphate $\overset{+2}{\text{Ca}}_3 (\overset{-3}{\text{PO}_4})_2$
- 26) ammonium sulfate $\overset{+1}{(\text{NH}_4)}_2 \overset{-2}{\text{SO}_4}$
- 27) silver cyanide $\overset{+1}{\text{Ag}} \overset{-1}{\text{CN}}$
- 28) aluminum sulfide $\overset{+3}{\text{Al}}_2 \overset{-2}{\text{S}_3}$
- 29) beryllium chloride $\overset{+2}{\text{Be}} \overset{-1}{\text{Cl}_2}$
- 30) copper (I) arsenide $\overset{+1}{\text{Cu}}_3 \overset{-3}{\text{As}}$
- 31) iron (III) oxide $\overset{+3}{\text{Fe}}_2 \overset{-2}{\text{O}_3}$
- 32) gallium nitride $\overset{+3}{\text{Ga}} \overset{-3}{\text{N}}$
- 33) iron (II) bromide $\overset{+2}{\text{Fe}} \overset{-1}{\text{Br}_2}$
- 34) vanadium (V) phosphate $\overset{+5}{\text{V}}_2 (\overset{-3}{\text{PO}_4})_5$
- 35) calcium oxide $\overset{+2}{\text{Ca}} \overset{-2}{\text{O}}$
- 36) magnesium acetate $\overset{+2}{\text{Mg}} (\overset{-1}{\text{C}_2\text{H}_3\text{O}_2})_2$
- 37) aluminum sulfate $\overset{+3}{\text{Al}}_2 (\overset{-2}{\text{SO}_4})_3$
- 38) copper (I) carbonate $\overset{+1}{\text{Cu}}_2 \overset{-2}{\text{CO}_3}$
- 39) barium oxide $\overset{+2}{\text{Ba}} \overset{-2}{\text{O}}$
- 40) ammonium sulfite $\overset{+1}{(\text{NH}_4)}_2 \overset{-2}{\text{SO}_3}$
- 41) silver bromide $\overset{+1}{\text{Ag}} \overset{-1}{\text{Br}}$
- 42) lead (IV) nitrite $\overset{+4}{\text{Pb}} (\overset{-1}{\text{NO}_2})_4$