

Name

KEK

Date:

Per:

Naming Ionic Compounds

Name the following Ionic Compounds:

- | | |
|---|---|
| 1. CuCl_2 <u>Copper (II) chloride</u> | 21. MgO <u>Magnesium oxide</u> |
| 2. CuCl <u>Copper (I) chloride</u> | 22. Na_2S <u>Sodium sulfide</u> |
| 3. FeCl_3 <u>Iron (III) chloride</u> | 23. SnCl_2 <u>Tin (II) chloride</u> |
| 4. FeO <u>Iron (II) oxide</u> | 24. SnCl_4 <u>Tin (IV) chloride</u> |
| 5. PbI_2 <u>Lead (II) Iodide</u> | 25. CuCl_2 <u>Copper (II) chloride</u> |
| 6. PbCl_4 <u>Lead (IV) chloride</u> | 26. BaI_2 <u>Barium Iodide</u> |
| 7. SnO <u>Tin (II) oxide</u> | 27. Al_2S_3 <u>Aluminum sulfide</u> |
| 8. SnO_2 <u>Tin (IV) oxide</u> | 28. LiI <u>Lithium Iodide</u> |
| 9. MnCl_6 <u>Manganese (VI) chloride</u> | 29. HgS <u>Mercury (II) sulfide</u> |
| 10. CoBr_3 <u>Cobalt (III) bromide</u> | 30. MgCl_2 <u>Magnesium chloride</u> |
| 11. MnF_3 <u>Manganese (III) fluoride</u> | 31. BaCl_2 <u>Barium chloride</u> |
| 12. Fe_2O_3 <u>Iron (III) oxide</u> | 32. Hg_2O <u>Mercury (I) oxide</u> |
| 13. Cu_2O <u>Copper (I) oxide</u> | 33. CuI <u>Copper (I) iodide</u> |
| 14. PbO <u>Lead (II) oxide</u> | 34. CaBr_2 <u>Calcium bromide</u> |
| 15. Na_2O <u>Sodium oxide</u> | 35. AgF <u>Silver fluoride</u> |
| 16. HgCl_2 <u>Mercury (II) chloride</u> | 36. PbCl_2 <u>Lead (II) chloride</u> |
| 17. K_2S <u>Potassium sulfide</u> | 37. HgBr <u>Mercury (I) bromide</u> |
| 18. CsBr <u>Cesium bromide</u> | 38. CoCl_2 <u>Cobalt (II) chloride</u> |
| 19. AlCl_3 <u>Aluminum chloride</u> | 39. SrS <u>Strontium sulfide</u> |
| 20. Fe_2O_3 <u>Iron (III) oxide</u> | 40. Ag_2O <u>Silver oxide</u> |

Note:

Do not forget to use roman numerals for the charge on the transition metals with more than one type of ion (some metals in groups 3 - 12).

One = I

Two = II

three = III

four = IV

five = V

six = VI

seven = VII