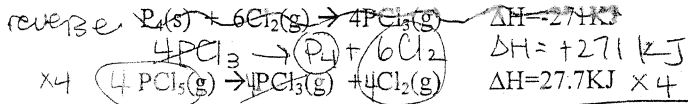
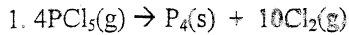
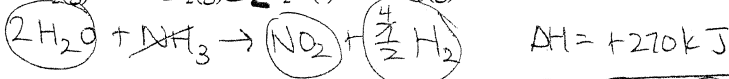
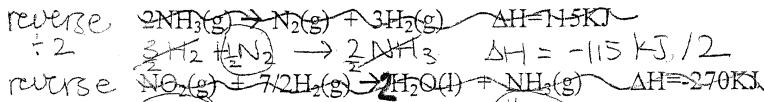
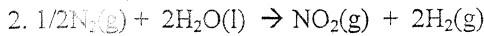


Hess's Law Worksheet

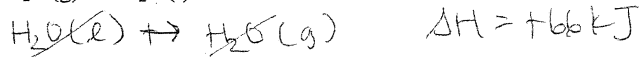
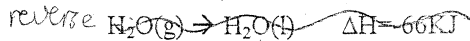
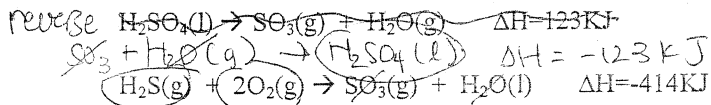
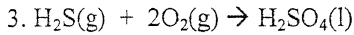
Directions: Use the thermochemical equations shown below each reaction to determine its enthalpy.



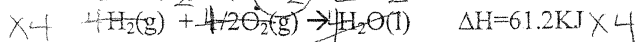
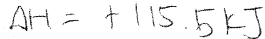
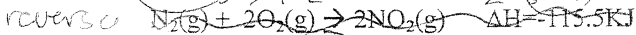
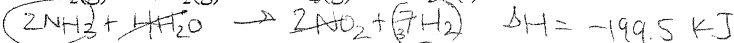
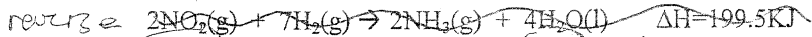
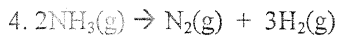
$$\Delta H = 381.8\text{KJ}$$



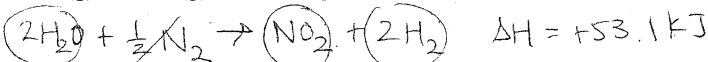
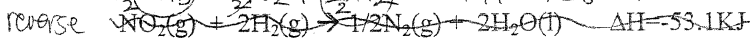
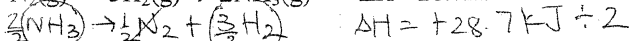
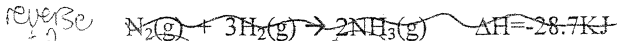
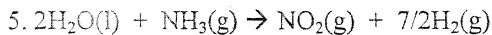
$$\Delta H = 212.5\text{KJ}$$



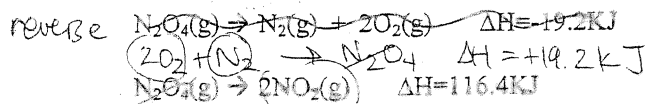
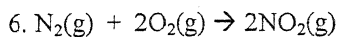
$$\Delta H = -471\text{KJ}$$



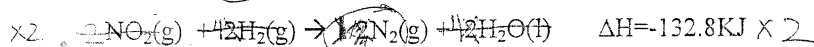
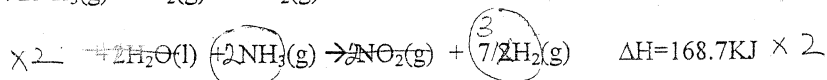
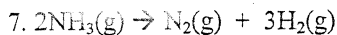
$$\Delta H = +160.8\text{KJ}$$



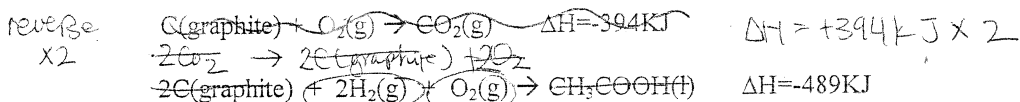
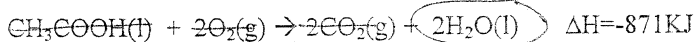
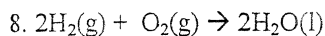
$$\Delta H = 67.45\text{KJ}$$



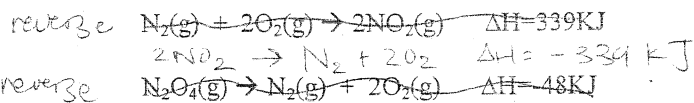
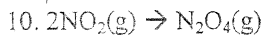
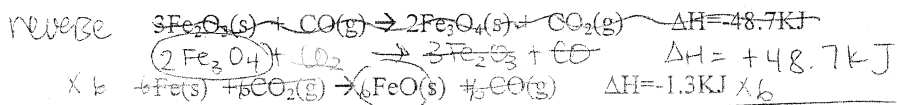
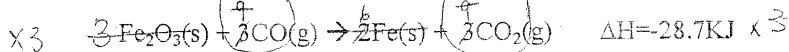
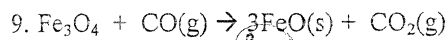
$\Delta H = 135.6 \text{ kJ}$



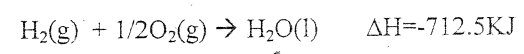
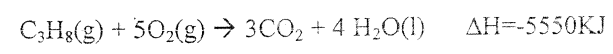
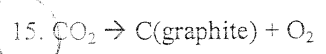
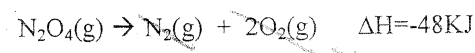
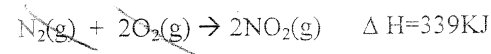
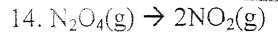
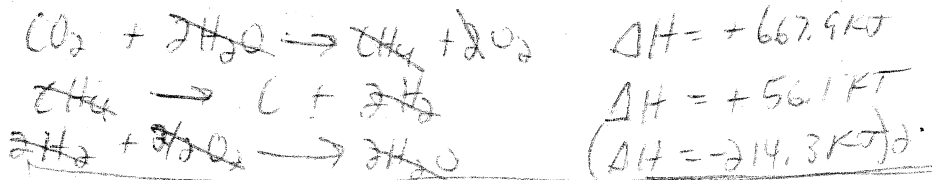
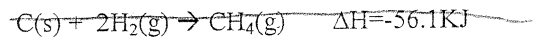
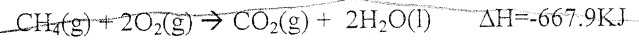
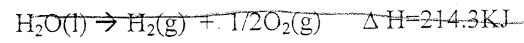
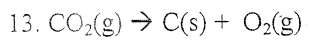
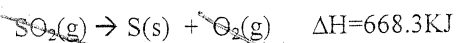
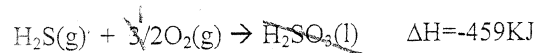
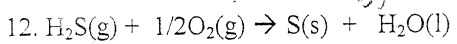
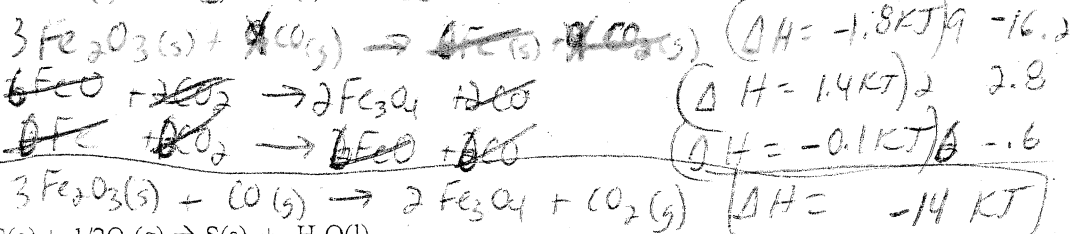
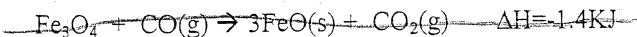
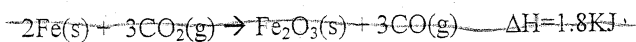
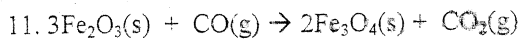
$\Delta H = 71.8 \text{ kJ}$



$\Delta H = -572 \text{ kJ}$



$\Delta H = -291 \text{ kJ}$



Missing C

