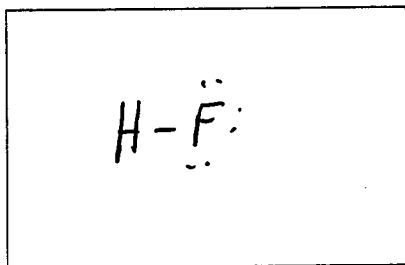


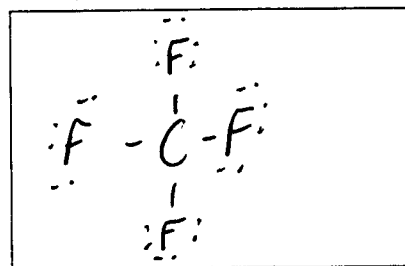


Give the Lewis Structure for each of the following:

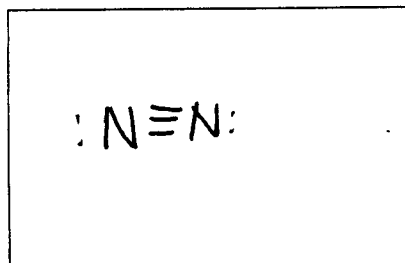
14. HF



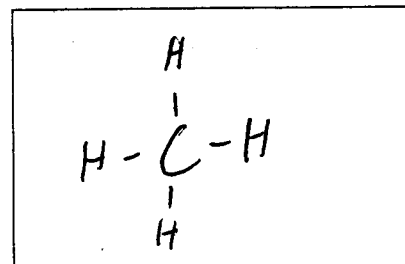
17. CF<sub>4</sub>



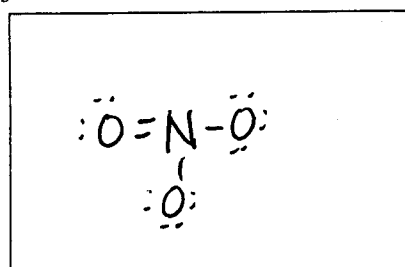
15. N<sub>2</sub>



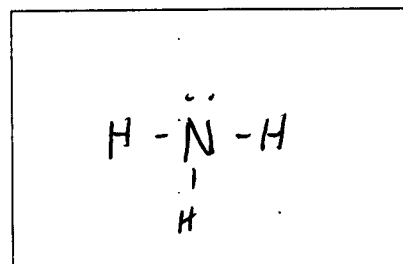
18. CH<sub>4</sub>



16. NO<sub>3</sub><sup>-1</sup>



19. NH<sub>3</sub>



On the basis of electronegativities, indicate whether each of the following bonds would be ionic, covalent, or polar covalent bonds: S=2.5, H=2.1, O=3.5, K=0.9

20. S-S Covalent

22. S-H Covalent

21. S-O polar covalent

23. S-K ionic or polar covalent

Give the total number of valence electrons in each of the following molecules.

24. CBr<sub>4</sub> 32

26.  $\overset{4}{\text{C}}\overset{1}{\text{H}}_6$  30

25.  $\overset{5}{\text{N}}\overset{6}{\text{O}}_2$  17

27.  $\overset{1}{\text{H}}_2\overset{6}{\text{O}}_2$  14