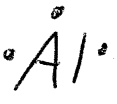
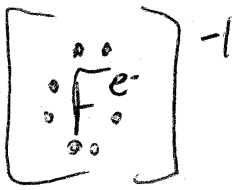
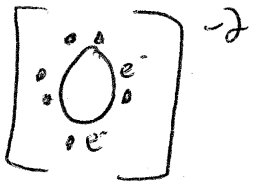
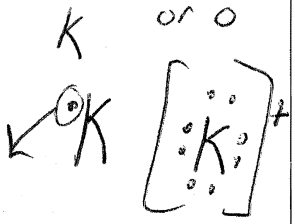
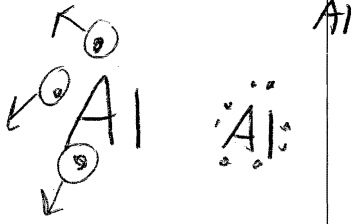
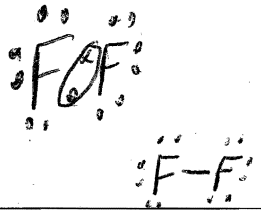
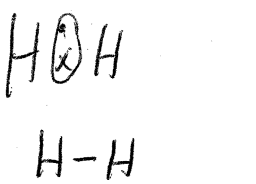
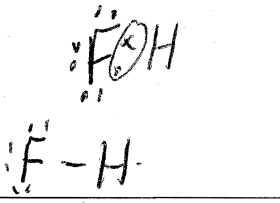
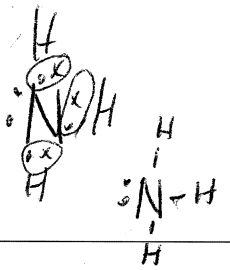
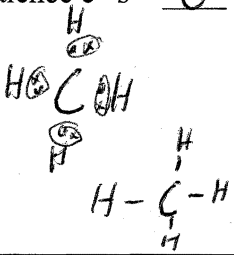
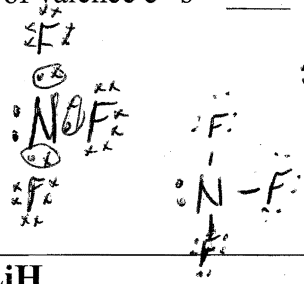
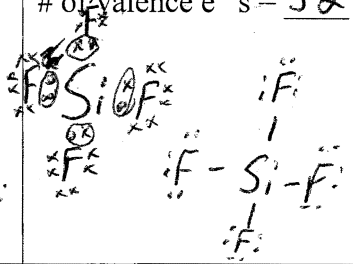
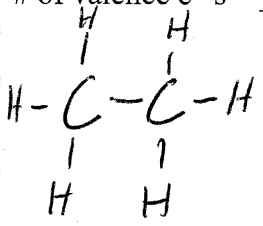
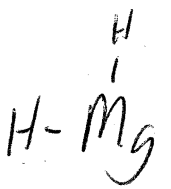
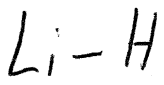
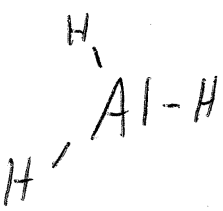
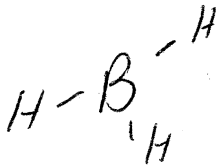
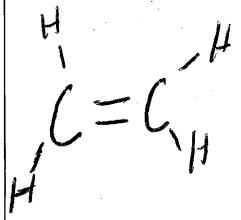
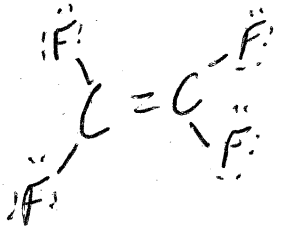
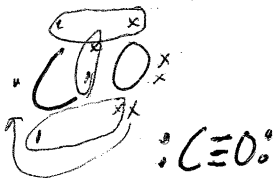
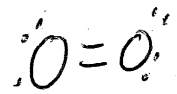
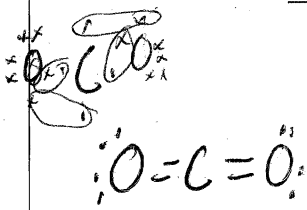
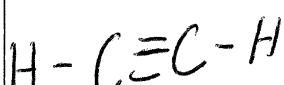
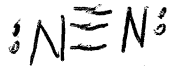
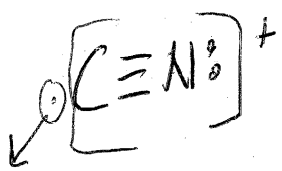
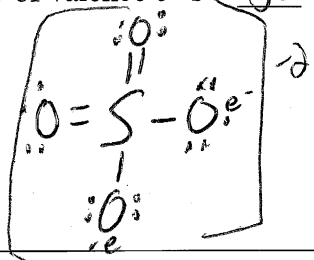
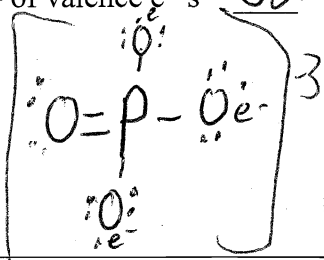
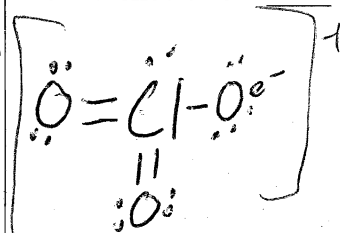
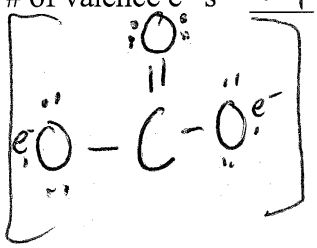
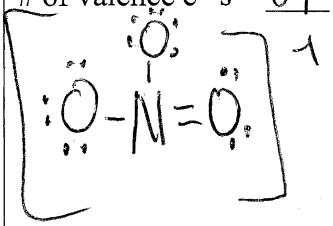
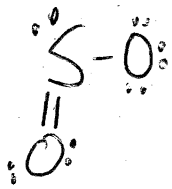
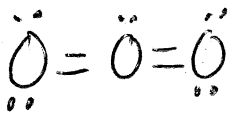
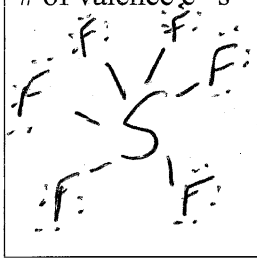
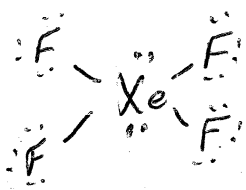
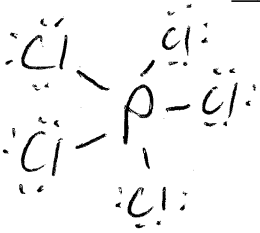
Name _____
Period ___ Date ___/___/___

6 • Bonding & Molecular Structure

LEWIS STRUCTURES

Indicate the # of **VALENCE** electrons for each species. Write the correct Lewis electron-dot structure for each.

F # of valence e⁻'s = 7 	**O** # of valence e⁻'s = 6 	**K** # of valence e⁻'s = 1 	**Al** # of valence e⁻'s = 3 
F^- # of valence e⁻'s = 8 	**O^{2-}** # of valence e⁻'s = 8 	**K^+** # of valence e⁻'s = 8 or 0 	**Al^{3+}** # of valence e⁻'s = 8 or 0 
F_2 # of valence e⁻'s = 14 	**H_2** # of valence e⁻'s = 2 	**HF** # of valence e⁻'s = 8 	**NH_3** # of valence e⁻'s = 8 
CH_4 # of valence e⁻'s = 8 	**NF_3** # of valence e⁻'s = ____ 	**SiF_4** # of valence e⁻'s = 32 	**C_2H_6** # of valence e⁻'s = 14 
MgH_2 # of valence e⁻'s = 4 	**LiH** # of valence e⁻'s = 2 	**AlH_3** # of valence e⁻'s = 6 	**BH_3** # of valence e⁻'s = 6 

C_2H_4 # of valence e⁻'s = 12 	**C_2F_4** # of valence e⁻'s = 36 	**CO** # of valence e⁻'s = 10 	**O_2** # of valence e⁻'s = 12 
CO_2 # of valence e⁻'s = 16 	**C_2H_2** (H C C H) # of valence e⁻'s = 10 	**N_2** # of valence e⁻'s = 10 	**HCN** # of valence e⁻'s = 10 
CN^- # of valence e⁻'s = 8 	**SO_4^{2-}** # of valence e⁻'s = 32 	**PO_4^{3-}** # of valence e⁻'s = 32 	**ClO_3^-** # of valence e⁻'s = 26 
CO_3^{2-} # of valence e⁻'s = 24 	**NO_3^-** # of valence e⁻'s = 24 	**~~SO_2~~** # of valence e⁻'s = 18 	**~~O_3~~** (O O O) # of valence e⁻'s = 18 
SF_6 # of valence e⁻'s = 48 	**XeF_4** # of valence e⁻'s = 38 	**PCl_5** # of valence e⁻'s = 40 	**SeF_4** # of valence e⁻'s = 34