

Name: KEY

Binary Molecular Writing & Naming

1) As_4O_{10}	<u>tetraarsenic decaoxide</u>	1) chlorine monoxide	<u>ClO</u>
2) BrO_3	<u>Bromine trioxide</u>	2) oxygen difluoride	<u>OF₂</u>
3) BN	<u>Boron mononitride</u>	3) boron phosphide	<u>BP</u>
4) N_2O_3	<u>dinitrogen trioxide</u>	4) dinitrogen monoxide	<u>N₂O</u>
5) Ni_3	<u>Nitrogen triiodide</u>	5) nitrogen trifluoride	<u>NF₃</u>
6) SF_6	<u>Sulfur hexafluoride</u>	6) sulfur tetrachloride	<u>SCl₄</u>
7) XeF_4	<u>Xenon tetrafluoride</u>	7) xenon trioxide	<u>XeO₃</u>
8) PCl_3	<u>Phosphorus trichloride</u>	8) carbon dioxide	<u>CO₂</u>
9) CO	<u>Carbon monoxide</u>	9) diphosphorous pentoxide	<u>P₂O₅</u>
10) PCl_5	<u>Phosphorus pentachloride</u>	10) phosphorous trichloride	<u>PCl₃</u>
11) P_2O_5	<u>diphosphorus pentoxide</u>	11) sulfur dioxide	<u>SO₂</u>
12) S_2Cl_2	<u>disulfur dichloride</u>	12) bromine pentafluoride	<u>BrF₅</u>
13) ICl_2	<u>Iodine dichloride</u>	13) disulfur dichloride	<u>S₂Cl₂</u>
14) SO_2	<u>Sulfur dioxide</u>	14) boron trifluoride	<u>BF₃</u>
15) P_4O_{10}	<u>tetraphosphorus decaoxide</u>	15) tetraarsenic decoxide	<u>As₄O₁₀</u>
16) UF_6	<u>Uranium hexafluoride</u>	16) silicon tetrachloride	<u>SiCl₄</u>
17) OF_2	<u>Oxygen difluoride</u>	17) krypton difluoride	<u>KrF₂</u>
18) ClO_2	<u>Chlorine dioxide</u>	18) chlorine monoxide	<u>ClO</u>
19) SiO_2	<u>Silicon dioxide</u>	19) silicon dioxide	<u>SiO₂</u>
20) BF_3	<u>Boron trifluoride</u>	20) boron trichloride	<u>BCl₃</u>
21) N_2S_5	<u>dinitrogen pentasulfide</u>	21) dinitrogen pentasulfide	<u>N₂S₅</u>
22) CO_2	<u>Carbon dioxide</u>	22) carbon monoxide	<u>CO</u>
23) SO_3	<u>Sulfur trioxide</u>	23) sulfur trioxide	<u>SO₃</u>
24) XeF_6	<u>Xenon hexafluoride</u>	24) dinitrogen trioxide	<u>N₂O₃</u>

Name KEY Date: _____ Per: _____

Naming Compounds – Ionic I, Ionic II and Covalent Compounds- HOMEWORK

DIRECTIONS: Name the following Ionic I, Ionic II, and Covalent Compounds:

1. CoCl_3 Cobalt(III) chloride
2. Hg_2O Mercury (I) oxide
3. SnCl_4 Tin (IV) chloride
4. SiCl_4 Silicon ~~the~~ tetrachloride
5. P_2Cl_4 diphosphorus tetrachloride
6. BCl_3 Boron trichloride
7. B_2H_6 diBoron hexahydride
8. Ca_3N_2 Calcium Nitride
9. Ag_2S Silver Sulfide
10. CBr_4 Carbon tetrabromide
11. CuCl_2 Copper (II) chloride
12. ClF Chlorine monofluoride
13. HCl hydrogen chloride
14. Ag_2O Silver oxide
15. Na_2S Sodium Sulfide
16. MgI_2 Magnesium Iodide
17. Al_2O_3 Aluminum oxide
18. BeO Beryllium oxide
19. CaH_2 Calcium hydride
20. LiF Lithium Fluoride
21. FeBr_2 Iron(II) bromide
22. CoS Cobalt (II) sulfide
23. SnO_2 Tin (IV) oxide
24. Co_2S_3 Cobalt(III) sulfide
25. HgCl_2 Mercury (II) chloride
26. Hg_2Cl_2 Mercury (I) chloride
27. PbI_4 Lead (IV) iodide
28. Fe_2O_3 Iron (III) oxide
29. CoBr_3 Cobalt (III) bromide
30. SnS_2 Tin(IV) sulfide
31. CoO Cobalt(II) oxide
32. CoBr_2 Cobalt (II) bromide
33. FeS Iron (II) sulfide
34. SnBr_4 Tin (IV) bromide
35. SnO Tin (II) oxide
36. OF_2 Oxygen difluoride
37. AsI_3 Arsenic trifiodide
38. XeF_6 Xenon hexafluoride
39. SF_6 Sulfur hexafluoride
40. N_2O_4 dinitrogen tetroxide
41. Cl_2O dichlorine monoxide

NOTE:

Ionic Compounds with metals that make one type of ion: metal name nonmetal-**ide**

Ionic Compounds with metals that make more than one type of ion: metal name (roman number) nonmetal-**ide**

Covalent Compounds with two different nonmetals: use prefixes and end the name of second nonmetal with -**ide**